

# **SNS COLLEGE OF PHARMACY AND HEALTH SCIENCES**

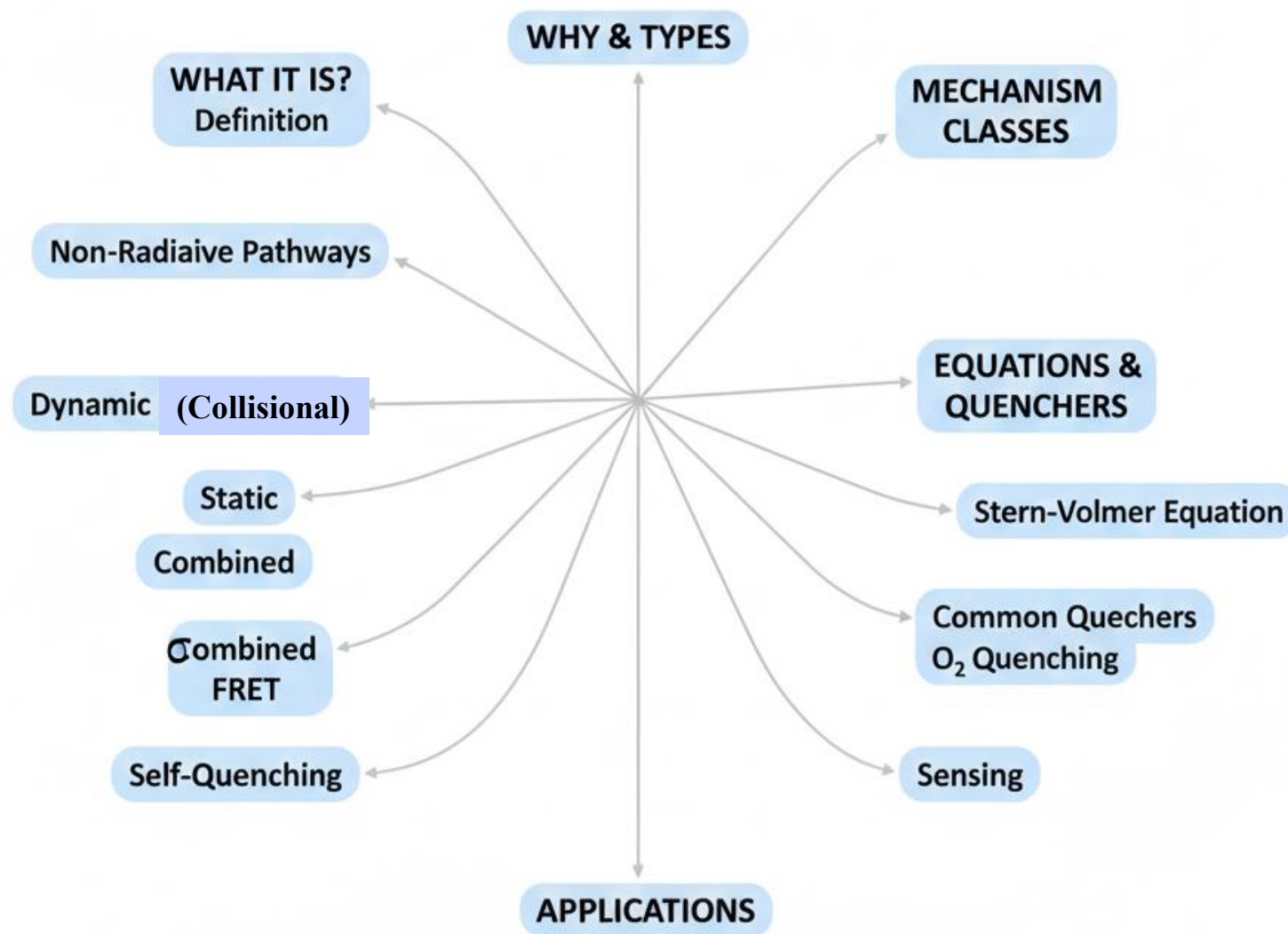
*Affiliated To The Tamil Nadu Dr. MGR Medical University, Chennai*  
*Approved by Pharmacy Council of India, New Delhi.*  
**Coimbatore -641035**

**COURSE NAME: INSTRUMENTAL METHODS OF ANALYSIS (BP 701 T)**

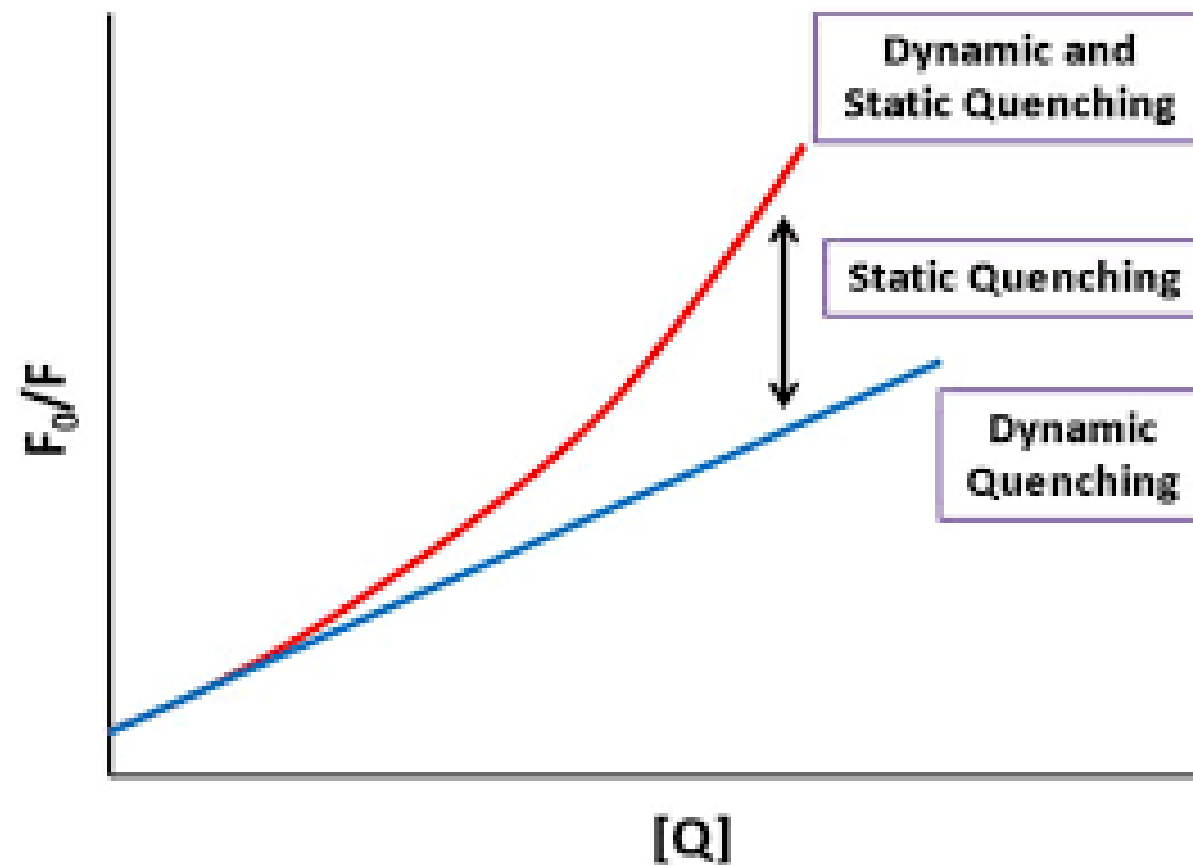
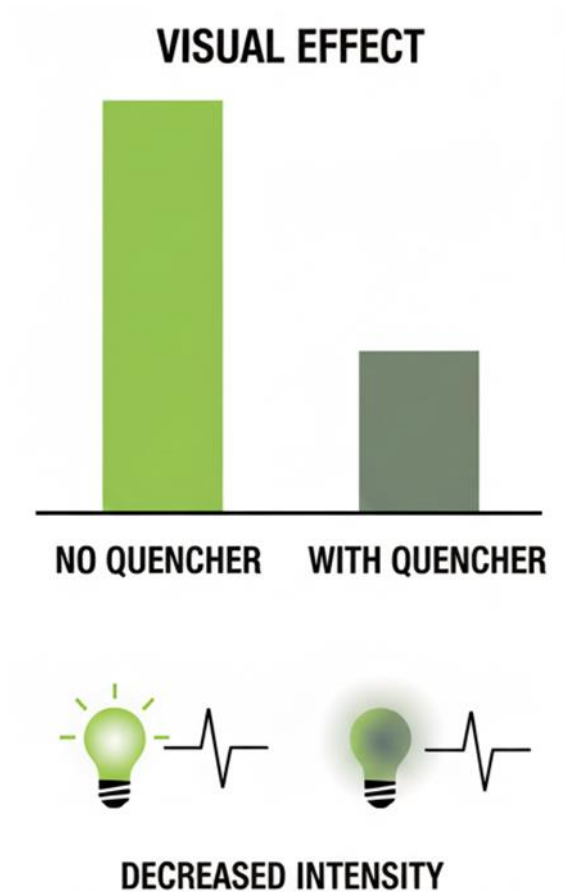
**VII SEM/ IV YEAR**

**TOPIC 9: QUENCHING**

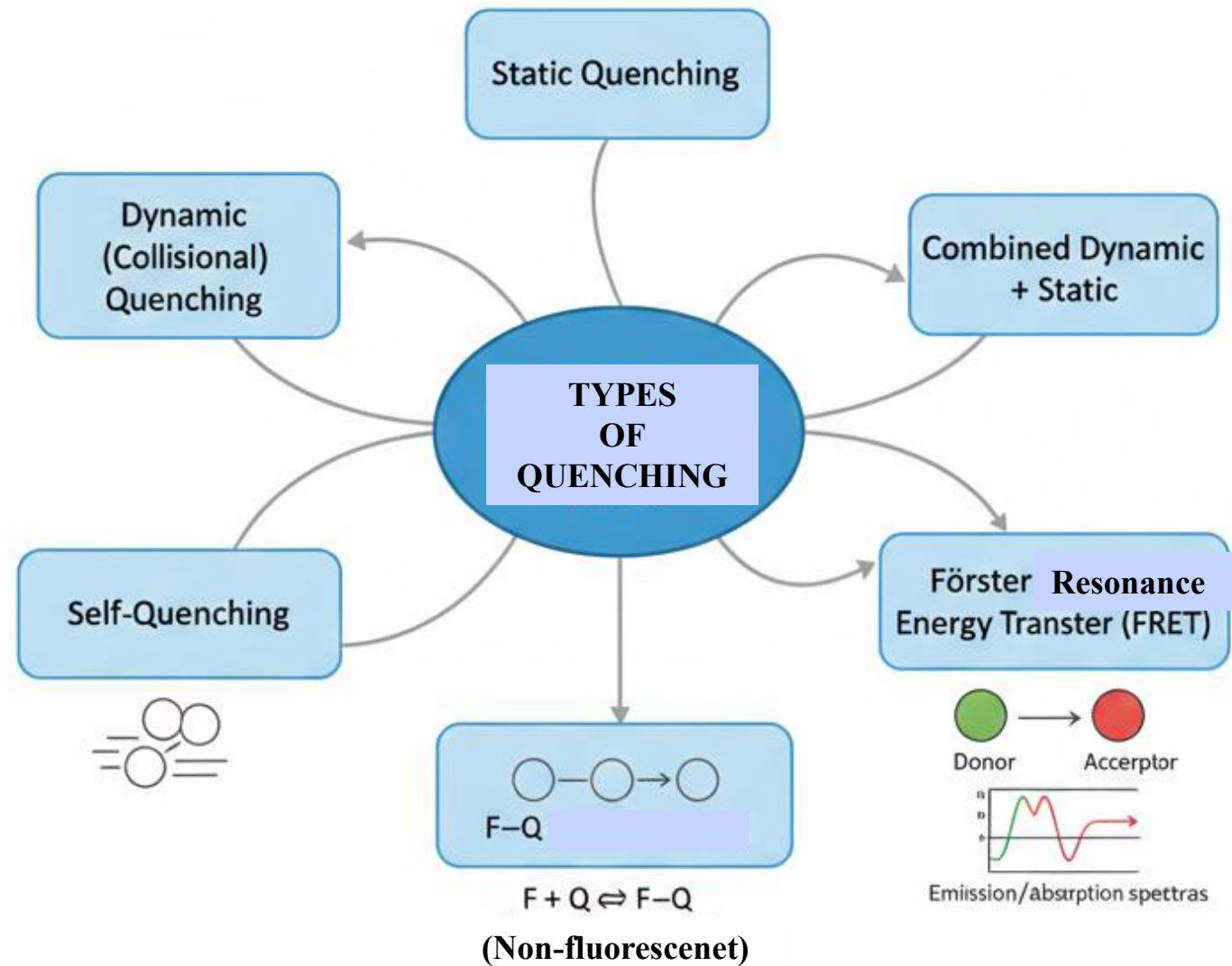
# MINDMAP:



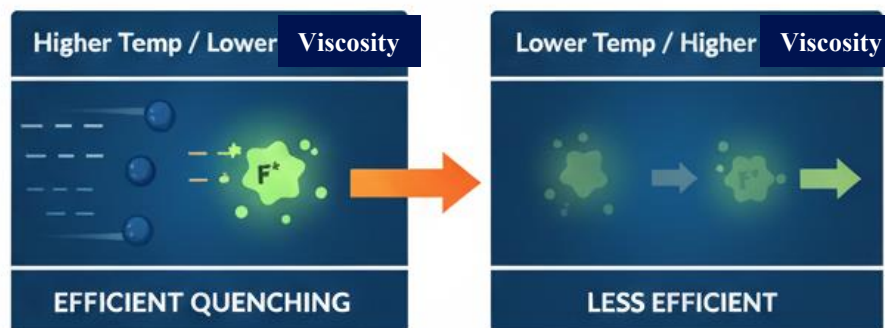
# Quenching



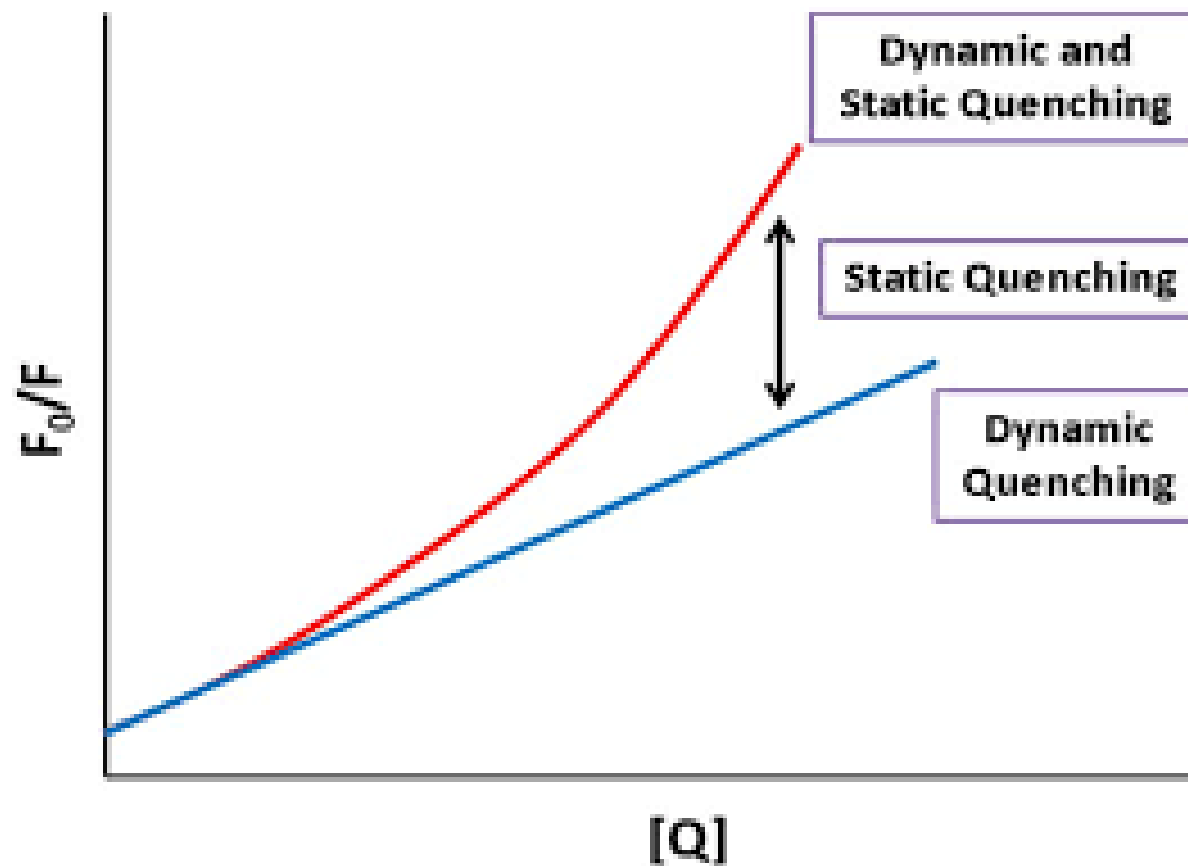
# Types of Quenching



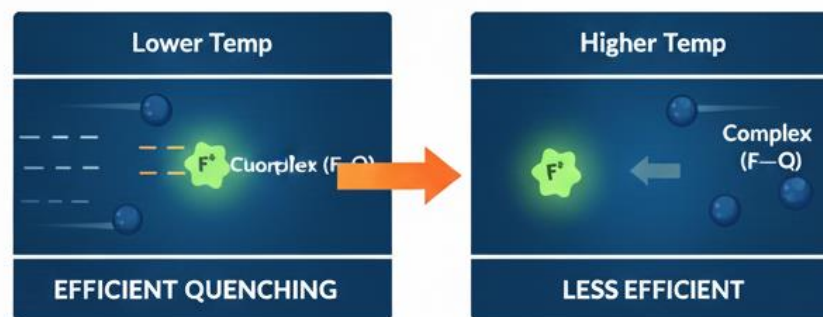
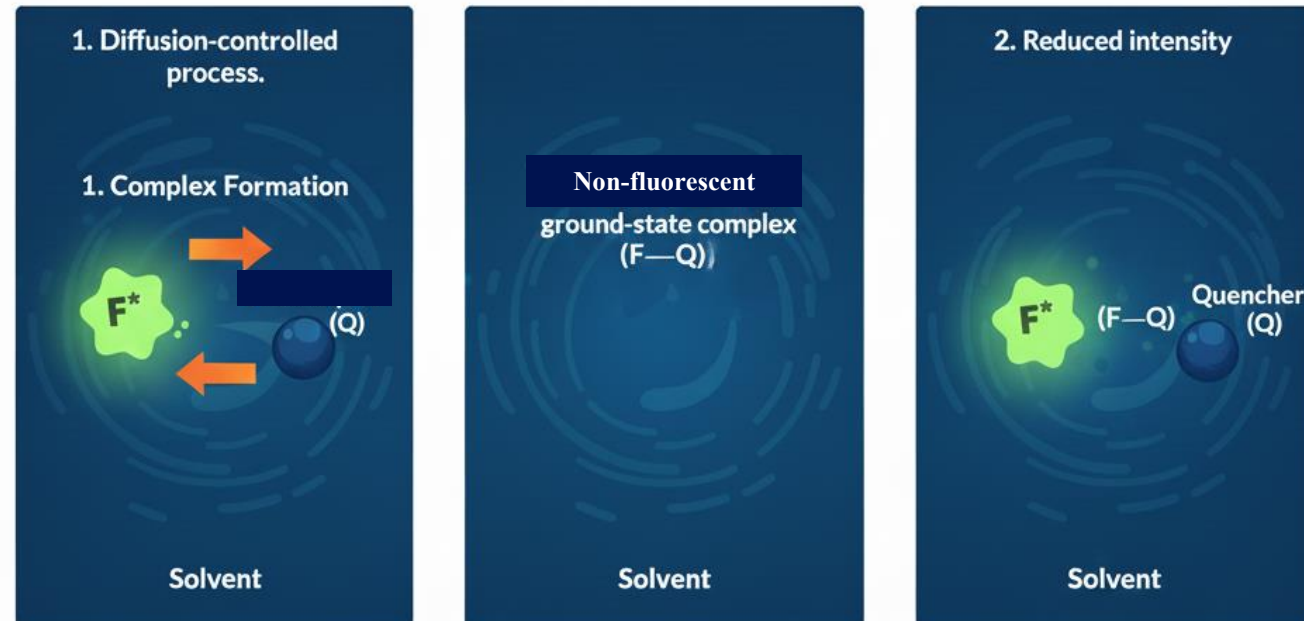
# Dynamic (Collisional) Quenching



# Stern-Volmer Equation – Dynamic Quenching

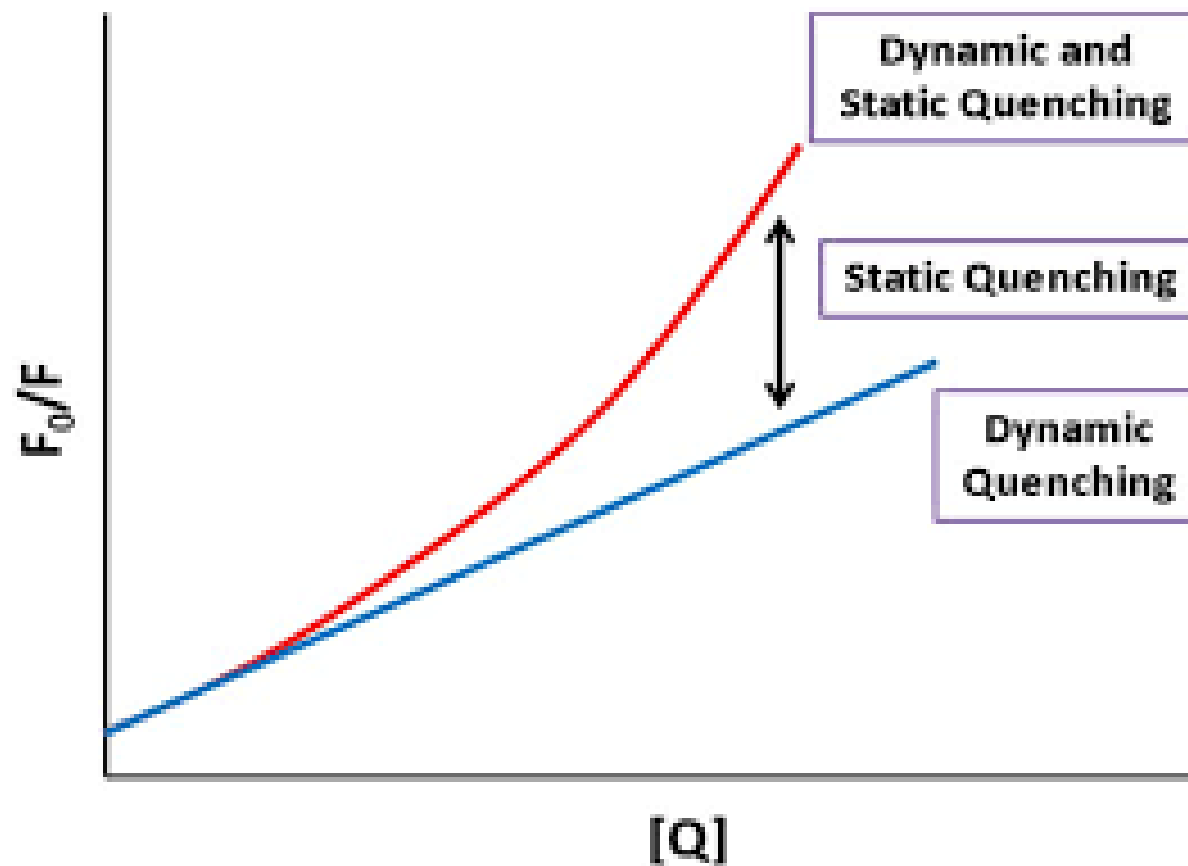


# Static Quenching



Reduces Intensity. No effect on Lifetime.

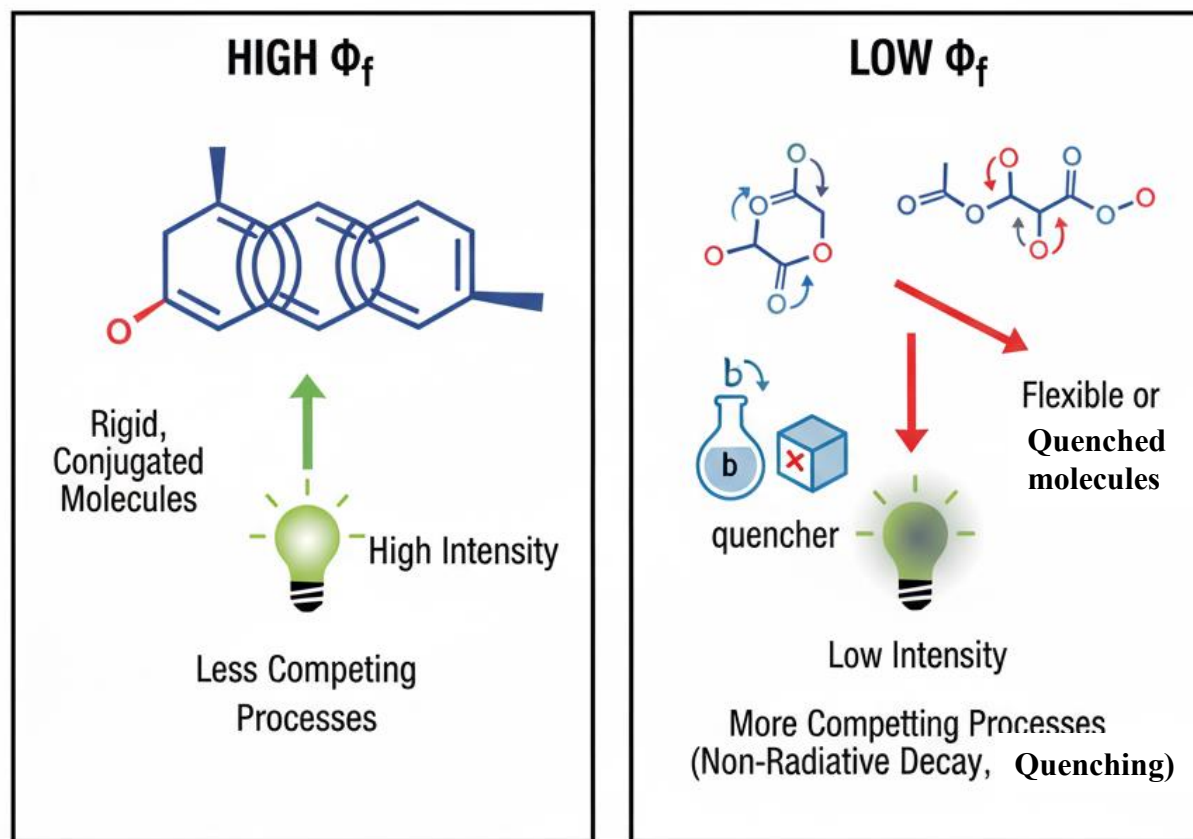
# Stern-Volmer Equation – Static Quenching





# Quantum Yield ( $\Phi_f$ )


$$\Phi_f = \text{Photons Emitted} / \text{Photons Absorbed}$$



Influenced by all competing processes (e.g., non-radiative decay, quenching)


# Common quenchers

### Gases / Small Molecules




strongest, universal,  
Paramagnetic

Oxygen ( $O_2$ )




TEMPO  
(stable radical)


### Halides / Amides



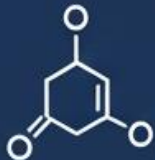
Iodide ( $I^-$ )




Bromide



$Br^-$




### Heavy Metal Ions




$+0$

$Cu^{2+}$



$Ni^{2+}$

$Ni^{2+}$

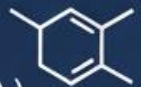


$+2^+$

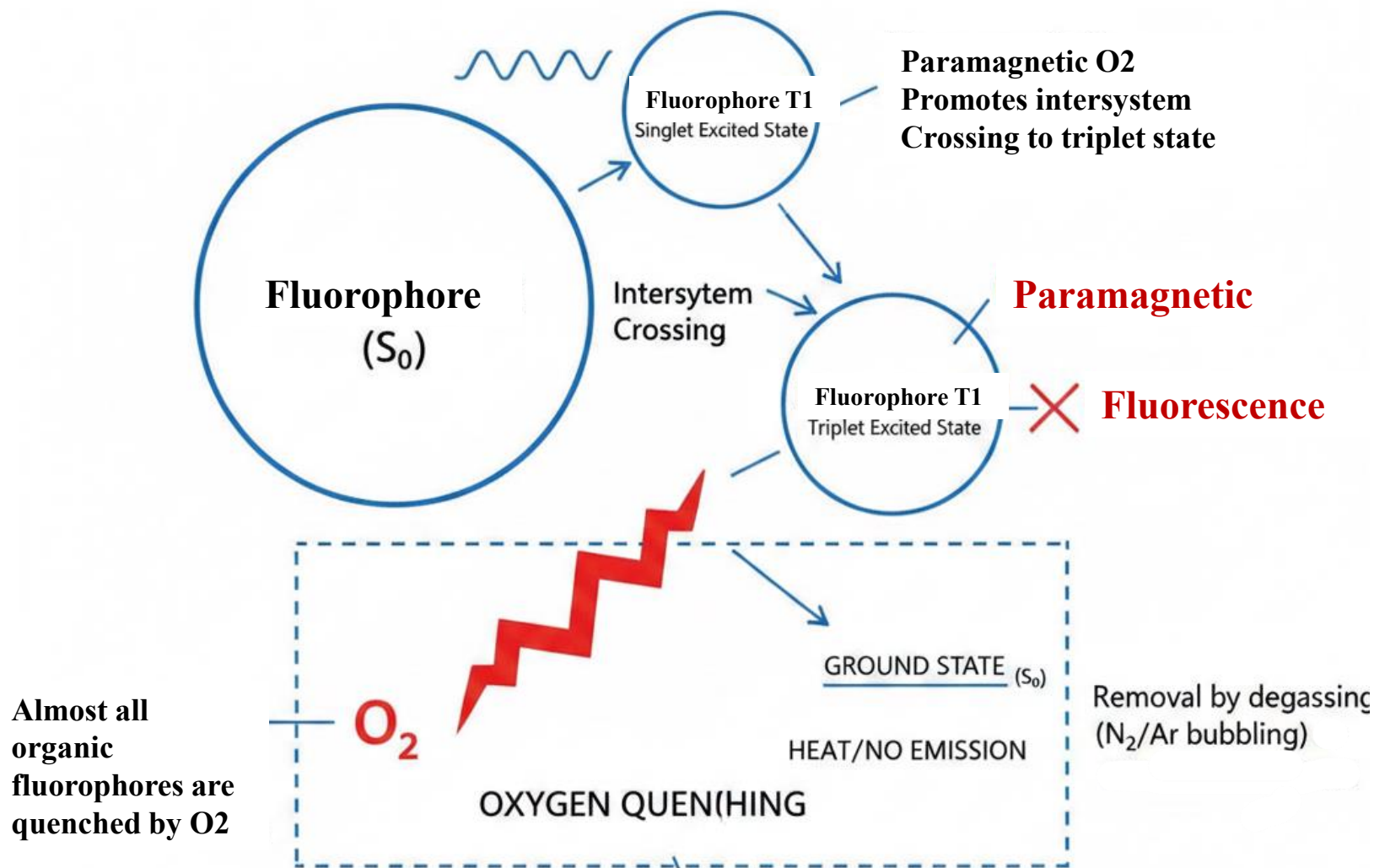
$Hg^{2+}$

### Nitromethane

( $CH_3NO_2$ )

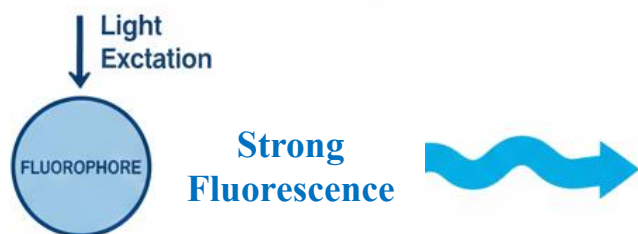


# Oxygen Quenching

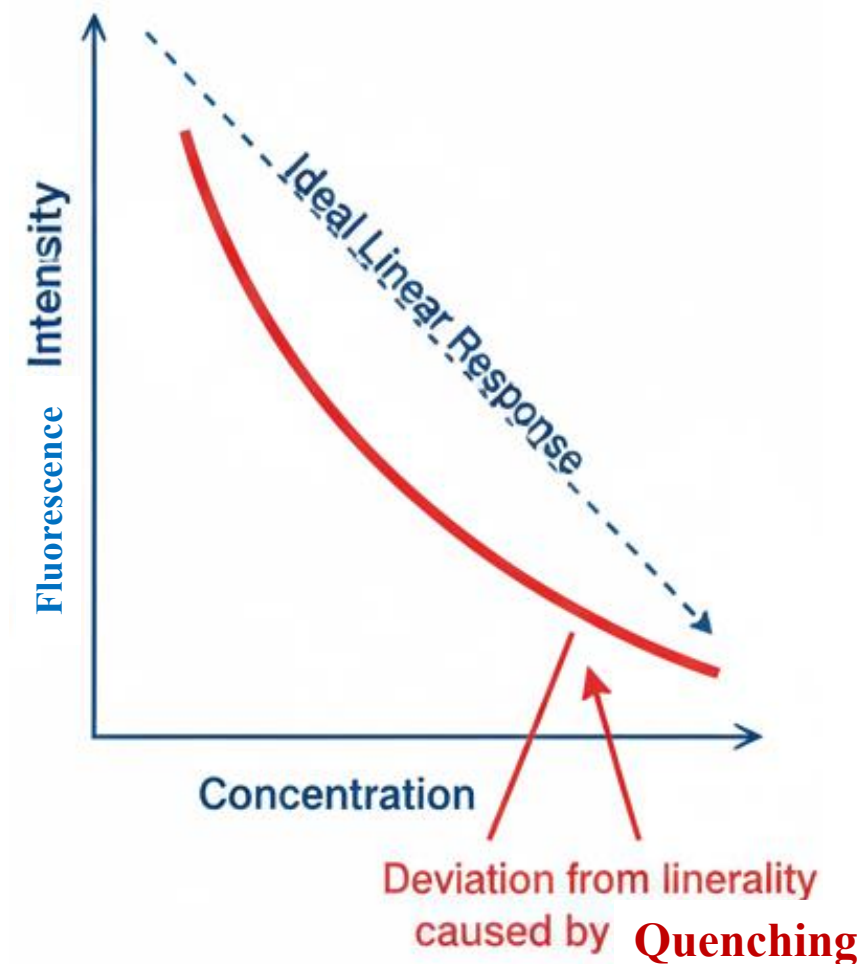
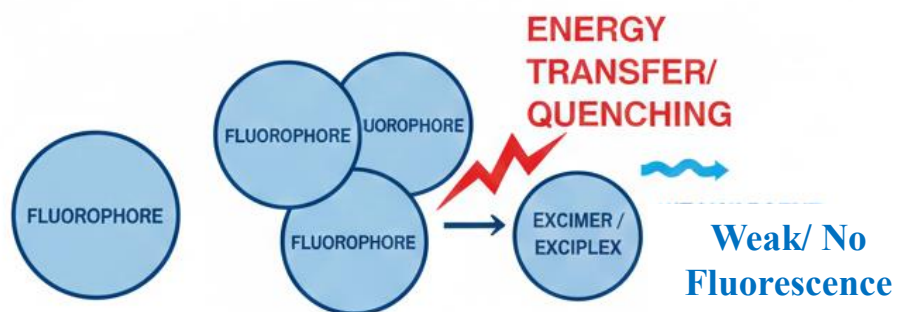


# Self/ concentration quenching

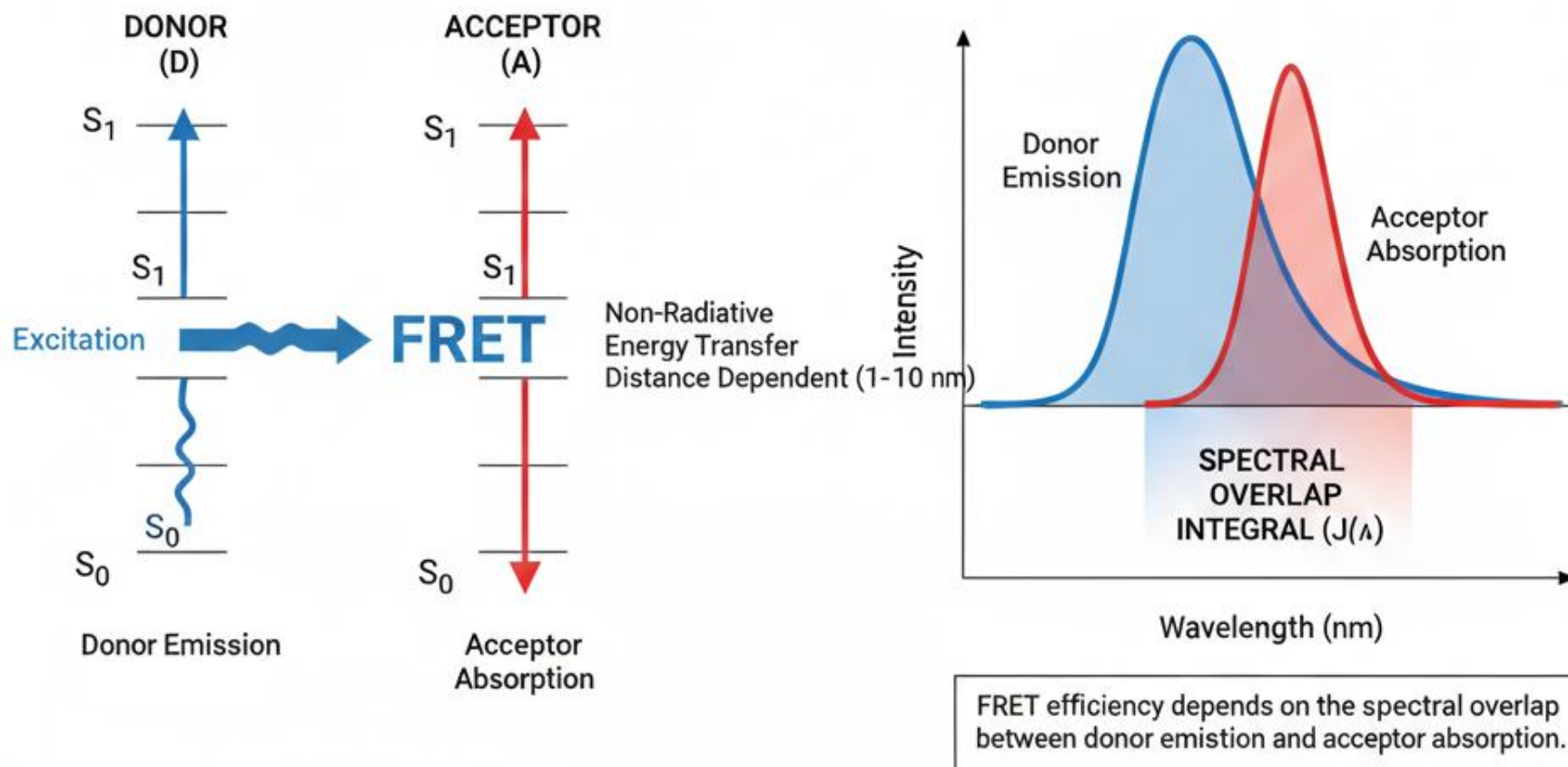
## Low Concentration



## High Concentration



# Förster Resonance Energy Transfer (FRET)



# Applications

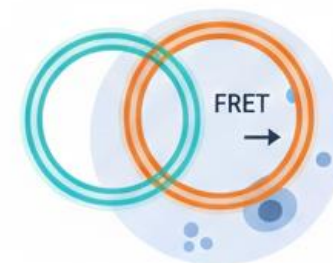
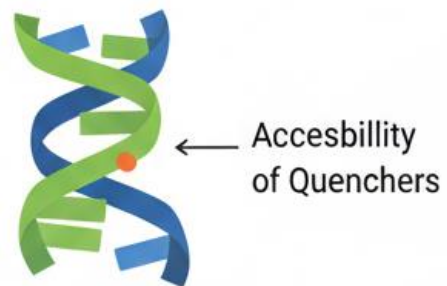
## 1. Determination of Quencher Concentration



## 2. Fluorescent based sensors



## 3. Protein conformation studies



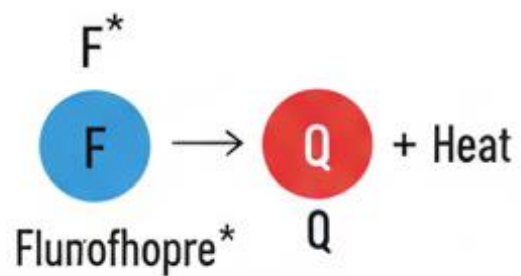
## 4. FRET in Biological Imaging

Biological Imaging & Assays

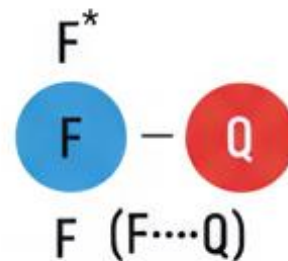


# Summary

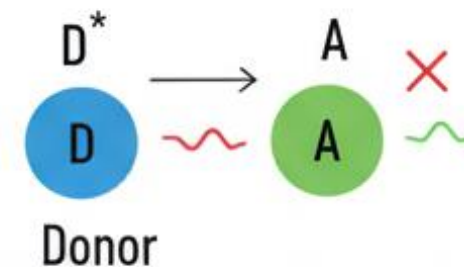
## 1. COLLISIONAL/DYNAMIC QUENCHING



## 2. STATIC QUENCHING



## 3. FRET/ENERGY TRANSFER



## Assessment

1. Fluorescence of phenol is stronger in:

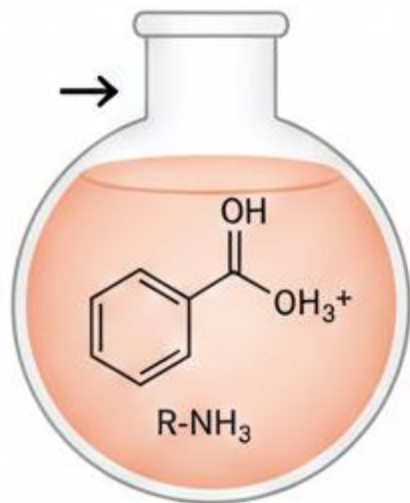




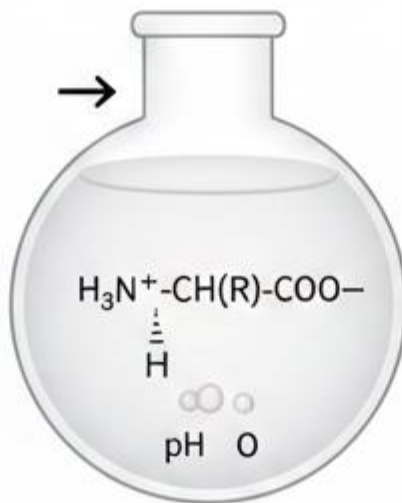
## Assessment

1. Fluorescence of phenol is stronger in:

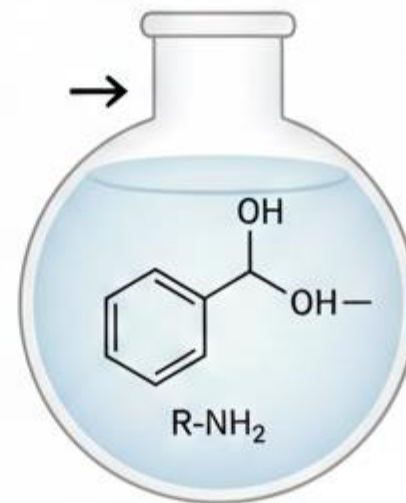
a) Acidic medium



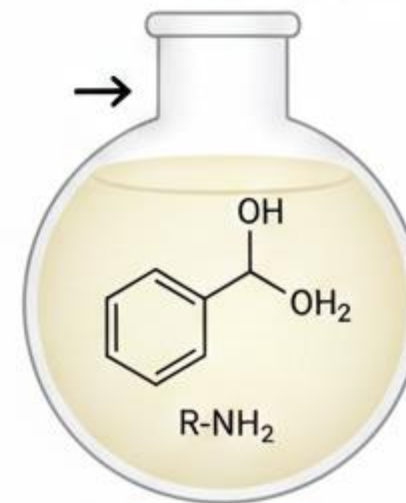
b) Neutral medium



c) Basic medium  
(ionized form)



d) Non-aqueous  
medium only



## Assessment

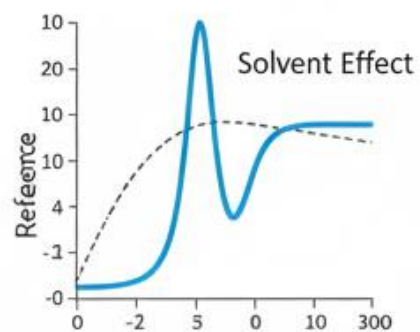
2. In polar protic solvents, the emission spectrum of a fluorophore often shows:



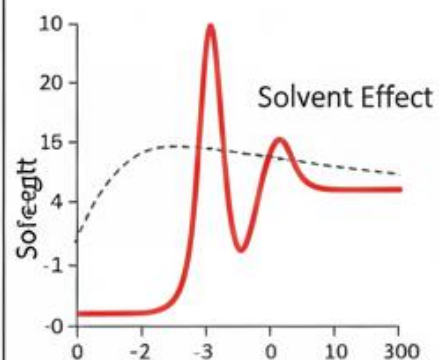
## Assessment

2. In polar protic solvents, the emission spectrum of a fluorophore often shows:

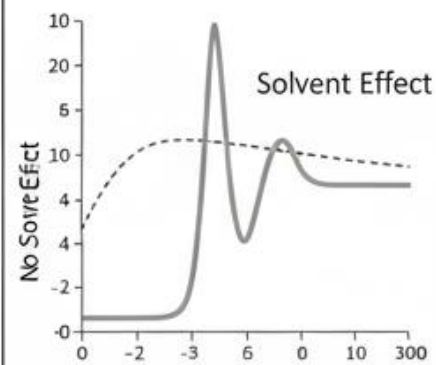
**a) Blue shift**



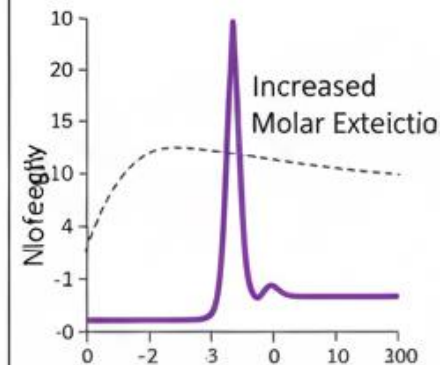
**b) Red shift due due solvent relaxation**



**c) No shift**



**d) Increased intensity only**



## Assessment

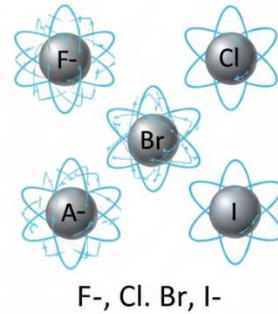
3. Which factor does NOT typically cause quenching?



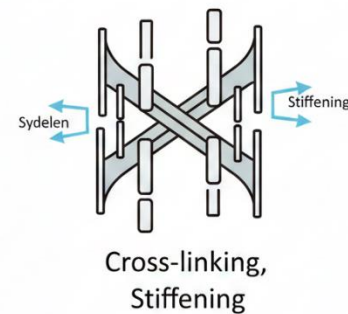
## Assessment

3. Which factor does NOT typically cause quenching?

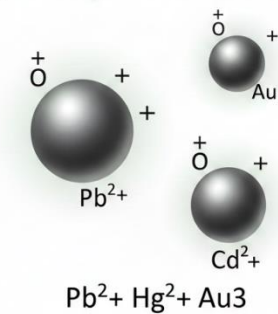
a) Halide ions



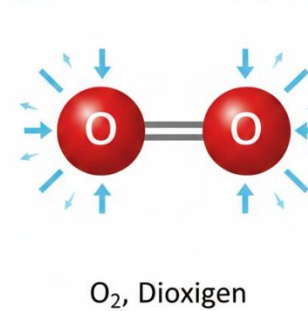
b) Increased rigidity



c) Heavy metal ions



d) Molecular oxygen



## References

1. Lakowicz JR. Principles of fluorescence spectroscopy. 3rd ed. New York (NY): Springer; 2006.
2. Skoog DA, Holler FJ, Crouch SR. Principles of instrumental analysis. 7th ed. Boston (MA): Cengage Learning; 2018.
3. Guilbault GG, editor. Practical fluorescence. 2nd ed. New York (NY): Marcel Dekker; 1990.
4. Valeur B, Berberan-Santos MN. Molecular fluorescence: principles and applications. 2nd ed. Weinheim (Germany): Wiley-VCH; 2012.

Thank  
you!